

Full Length Research Paper

Extraction and separation of Ni(II) from by functionalized graphene oxide with covalently linked porphyrin (GO-H₂NP) adsorbed on surfactant coated C₁₈

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A novel, simple, sensitive and effective method has been developed for preconcentration of Ni (II) on functionalized Graphene oxide with covalently linked porphyrin (GO-H₂NP) solid-phase extraction adsorbent. The stability of a chemical (GO-H₂NP), especially in concentrated hydrochloric acid, was then used as a recycling and pre-concentration reagent. The method is based on selective chelation of Ni(II) on surfactant coated C₁₈, modified with functionalized GO with covalently linked porphyrin (GO-H₂NP). The retained ions were then eluted with 4 ml of 4 M nitric acid and determined by flame atomic absorption spectrometry (FAAS) at 283.3 nm for Ni. The influence of flow rates of sample and eluent solutions, pH, breakthrough volume, effect of foreign ions on chelation and recovery were investigated. 1.5 g of surfactant coated C₁₈ adsorbs 40 mg of the functionalized Graphene oxide with covalently linked porphyrin (GO-H₂NP) base which in turn can retain 20.0 ± 0.9 mg of each of the two ions. The limit of detection (3σ) for Ni (II) was found to be 5.40 ngL⁻¹. The enrichment factor for both ions is 100. The mentioned method was successfully applied on the determination of Ni in different water samples. The ions were also speciated by means of a three columns system.

Key words: Determination of Ni, pre-concentration, functionalized graphene oxide, covalently linked porphyrin (GO-H₂NP), flame atomic absorption spectrometry (FAAS).

INTRODUCTION

Ni at trace concentrations acts as both a micronutrient and a toxicant in marine and fresh water systems (Izatt et al., 1985, 1991, 1995; Blake et al., 1996; Arca et al., 2001; Ghoulipour et al., 2002; Hashemi et al., 2001; Shcherbinina et al., 1990). This element is needed by plants at only very low levels and is toxic at higher levels. At these levels, Ni can bind to the cell membrane and hinder the transport process through the cell wall. Ni at

nearly 40 ng mL⁻¹ is required for normal metabolism of many living organisms (Gomes-Gomes, 1995; Moghimi et al., 2006; Unger et al., 1979).

On the other hand, Ni is an important element in many industries. Thus, the development of new methods for selective separation, concentration and determination of it in sub-micro levels in different industrial, medicinal and environmental samples is of continuing interest.

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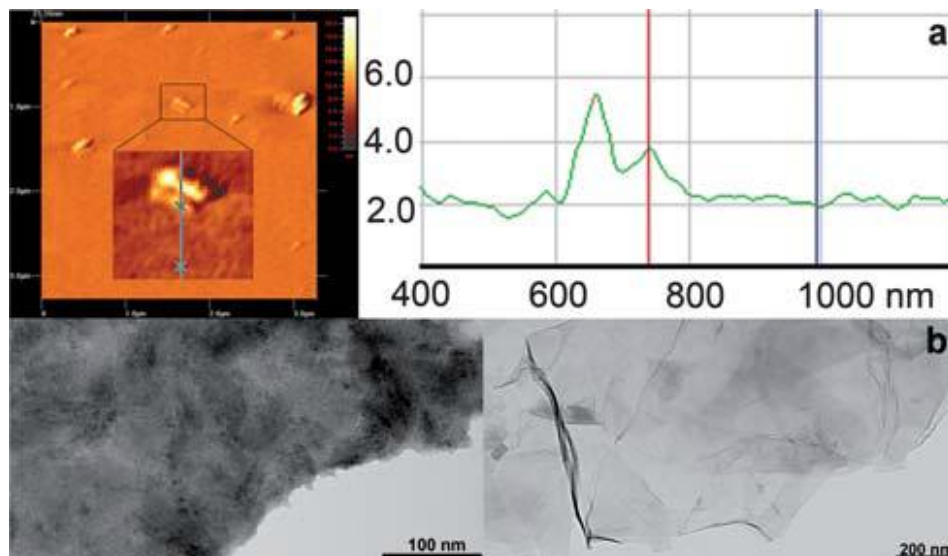


Figure 1. (a) Representative AFM image of GO-H₂NP and profile analysis showing a height of 1.77 nm for the enlarged region. Section analysis of other regions of the image show height ranges of 1.5–3.5 nm. (b) TEM images of the intact graphite (left panel) and GO-H₂NP hybrid material (right panel).

The determination of Ni is usually carried out by flame and graphite furnace atomic absorption spectrometry (AAS) (Boudreau et al., 1989) as well as spectrometric methods (Bruening et al., 1991; Mahmoud, 1997a; Mahmoud and Soliman, 1997).

Solid phase extraction (SPE) methods are the best alternatives for traditional classic methods due to selective removal of trace amounts of metal ions from their matrices. Solid phase extraction determinations can be carried out on different efficient ways. One of the most appropriate performance features of SPE is achieved by using octadecyl silica membrane disks. SPE reduce the use of toxic solvent, disposal costs, and extraction time (Mahmoud, 1997b; Talata, 1997:45-44). The octadecyl silica membrane disks involves shorter sample processing time and decreased plugging due to the large cross-sectional area of the disk and small pressure drop which allows higher flow-rates; reduced channeling resulting from the use of sorbent with smaller particle size and a greater mechanical stability of the sorbent bed (Tong et al., 1990).

In our previous attempts, we modified SPE membrane disks with suitable compounds for selective determination of chromium (Dadler et al., 1987; Moghimi, 2007) and lead (Mahmoud 1999). Meanwhile, other investigators have successfully utilized these sorbents for quantitative extraction and monitoring trace amounts of lead (Leyden et al., 1976; Moghimi et al., 2009; Liu et al., 1992), copper (Liu et al., 1996; Mishenina et al., 1996; Wang et al., 1999), silver (Wang et al., 1997; Zhang et al., 1982), mercury (Zhou et al., 1983; Zargaran et al., 2008), cadmium (Tabarzadi et al., 2010), palladium (Shin et al.,

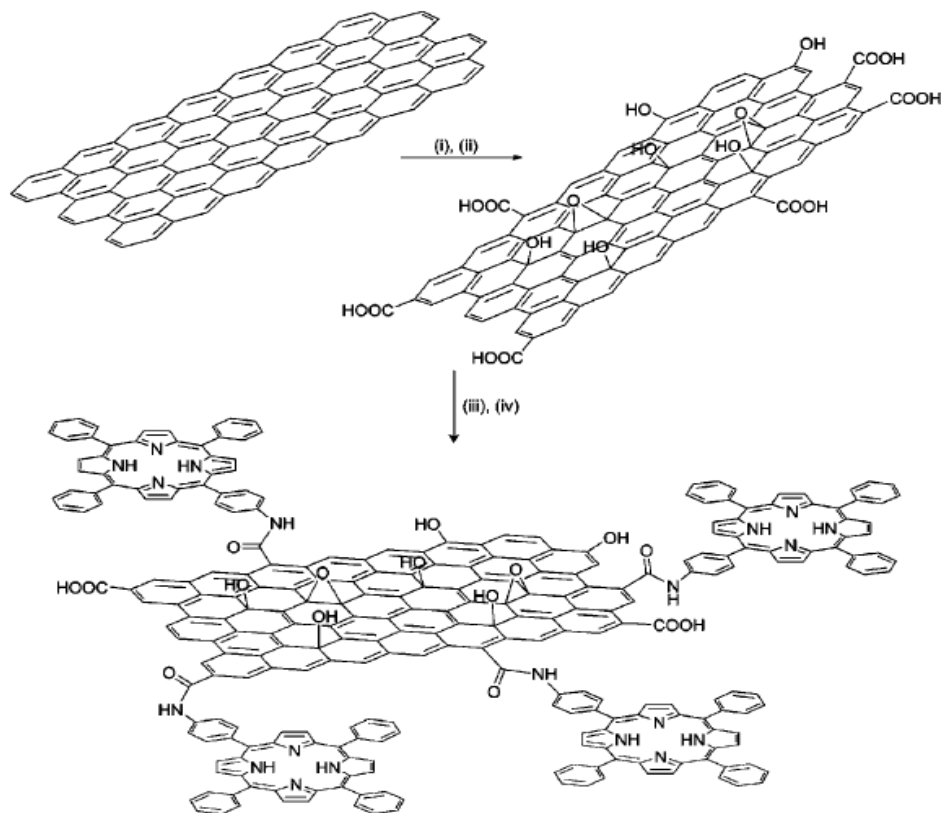
2004), Ce (Nayebi et al., 2006) and UO₂ (Mahmoud 1998). The main goal of the present work is the development of a fast, sensitive and efficient way for enrichment and extraction of trace amounts of Ni (II) from aqueous media by means of a surfactant coated C₁₈ modified with GO with covalently linked porphyrin is shown in Figure 1. Such a determination has not been reported in the literature.

The structure of functionalized Graphene oxide with covalently linked porphyrin (GO-H₂NP) is shown in Figure 1. The chelated ions were desorbed and determined by FAAS. The modified solid phase could be used at least 50 times with acceptable reproducibility without any change in the composition of the sorbent, GO-H₂NP or SDS. This work investigated the applications of nano Graphene oxide with covalently linked porphyrin (GO-H₂NP) like the C₁₈ SPE mini-column for selective extraction and solid phase pre-concentration of Ni (II) from aqueous and natural water samples.

EXPERIMENTAL

Reagents and apparatus

All solutions were prepared with doubly distilled deionized water. C₁₈ powder for chromatography with a diameter of about 50 μm was obtained from Katayama Chemicals. It was conditioned before use by being suspended in 4 M nitric acid for 20 min, and then washed two times with water. Sodium dodecyl sulfate (SDS) was obtained from Merck and used without any further purification. 1-(3-Dimethylaminopropyl)-3-ethylcarbodiimide hydrochloride (EDC) and N-hydroxysuccinimide (NHS) were purchased from Aldrich.



Schematic 1. A schematic illustration for the preparation of GO with covalently linked H₂NP. (i) H₂SO₄/HNO₃ (2 : 1 v/v), (ii) KClO₃, 96 h, (iii) (COCl)₂, 80 °C, 24 h, (iv) 5-(4-aminophenyl)-10,15,20-triphenyl-21,23H-porphyrin, THF, r.t., 72 h.

Graphene oxide (Guangfu Fine Chemical Research Institute, Tianjin, China) was used to prepare GO–H₂NP.

Preparation of GO

Graphite (5 g) was added to a vigorously stirring, ice-cooled mixture of sulfuric acid (88 ml) and nitric acid (44 ml). After the graphite was well dispersed, potassium chlorate (55 g) was added slowly, in portions, over 15 min, with simultaneous purging with nitrogen gas, to avoid a rapid increase of the temperature and to reduce the concentration of the generated chlorine dioxide gas. Then, the reaction mixture was allowed to stir for 96 h at room temperature. On completion of the reaction, the mixture was poured into 4 L of ice water and filtered through a 0.45 mm PTFE membrane filter. The graphite oxide was redispersed and washed with 5 L of a 5% HCl solution and then washed repeatedly with deionized water until the pH of the filtrate was neutral. The filter cake was then redispersed in methanol and after filtration was dried in a vacuum oven at 60°C overnight and then under atmospheric pressure at 120°C for 8 h to give GO as a browngray powder (3.9 g).

Synthesis of GO–H₂NP

GO (15 mg) was stirred in 20 ml of oxalyl chloride at 80°C for 24 h to activate the carboxylic units by forming the corresponding acyl chlorides. Then, the reaction mixture was evaporated to remove the

excess oxalyl chloride and the brownish remaining solid (GO–COCl) was washed with anhydrous tetrahydrofuran (THF). After centrifugation, the resulting solid material was dried at room temperature under vacuum. For the covalent coupling between the free amino function of H₂NP and the acyl chloride of GO, 15 mg of GO–COCl was treated under anaerobic, dry conditions with 7 mg of porphyrin dissolved in 6 ml of dry THF at room temperature for 72 h. The hybrid material, namely GO–H₂NP, was obtained as a brown-gray solid by filtration of the reaction mixture through 0.2 mm PTFE filter and the filtrate was sufficiently washed with methylene chloride (4 × 20 ml) to remove non-reacted free H₂NP and then with diethyl ether (2 × 20 ml) before being dried under vacuum (Schematic 1).

Column preparation

GO–H₂NP (40 mg) was packed into an SPE mini-column (6.0 cm × 9 mm i.d., polypropylene). A polypropylene frit was placed at each end of the column to prevent loss of the adsorbent. Before use, 0.5 mol L⁻¹ HNO₃ and DDW were passed through the column to clean it.

Apparatus

The pH measurements were conducted with an ATC pH meter (EDT instruments, GP 353) calibrated against two standard buffer

Table 1. The operational conditions of flame for determination of Ni²⁺.

Slit width	0.7 nm
Operation current of HI-HCL	10 mA
Resonance fine	283.3
Type of background correction	Deuterium lamp
Type of flame	Air/acetylene
Air flow	7.0 mL.min ⁻¹
Acetylene flow	1.7 mL.min ⁻¹

solutions of pH 4.0 and 9.2. Infrared spectra of GO–H₂NP were carried out from the KBr pellet by a Perkin-Elmer 1430 ratio recording spectrophotometer. Atomic absorption analysis of all the metal ions except Zn (II) was performed with a Perkin-Elmer 2380 flame atomic absorption spectrometer. Zn(II) determinations were performed by a Varian Spect AA-10. Raman spectrophotometer analysis was performed with a Perkin-Elmer.

Preparation of admicell column

Solution was added to 40 ml of water containing 1.5 g of C₁₈, and 150 mg of the above Graphene oxide with covalently linked porphyrin was loaded after washing acetone, 1 mol L⁻¹ HNO₃ solution and water, respectively. The pH of the suspension was adjusted to 2.0 with the addition of 4 mol L⁻¹ HNO₃ and stirred by a mechanical stirrer for 20 min.

Then the top liquid was decanted (and discarded) and the remaining C₁₈ was washed three times with water, then with 5 ml of 4 mol L⁻¹ HNO₃ and again three times with water. The prepared sorbent was transferred to a polypropylene tube (i.d 5 mm, length 10 mm).

Determination of Ni²⁺ contents in working samples were carried out by a Varian spectra A.200 model atomic absorption spectrometer equipped with a high intensity hollow cathode lamp (HI-HCl) according to the recommendations of the manufacturers. These characteristics are tabulated in Table 1. A metrohm 691 pH meter equipped with a combined glass calomel electrode was used for pH measurements.

Procedure

The pH of a solution containing 100 ng of each Ni(II) was adjusted to 2.0. This solution was passed through the admicell column with a flow rate of 5 ml min⁻¹. The column was washed with 10 ml of water and the retained ions were desorbed with 1 ml of 4 M HNO₃ with a flow rate of 2 ml min⁻¹. The desorption procedure was repeated 3 more times. All the acid solutions (4 ml all together) were collected in a 10 ml volumetric flask and diluted to the mark with water. The concentrations of Ni in the solution were determined by FAAS at 283.3.

Determination of Ni in water samples

Polyethylene bottles were soaked in 1 mol L⁻¹ HNO₃ overnight, and washed two times with water and used for sampling. The water sample was filtered through a 0.45 µm porous filter. The pH of a 1000 ml portion of each sample was adjusted to 2.0 (4 M HNO₃) and passed through the column under a flow rate of 5 ml min⁻¹. The column was washed with water and the ions were desorbed and determined as the above mentioned procedure.

Speciation of Ni in water samples

This procedure is reported in several articles. The method has been evaluated and optimized for speciation and its application on complex mixtures (Tabarzadi et al., 2010; Moghimi, 2007). The chelating cation exchanger (Chelex-100) and anion exchanger, (Dowex 1X-8) resins were washed with 1 M HCl, water, 1 M NaOH and water, respectively. 1.2 g of each resin was transferred to separate polyethylene columns. Each column was washed with 10 ml of 2 M HNO₃ and then 30 ml of water. The C₁₈ bounded silica adsorber was conditioned with 5 ml of methanol in a separate column, then with 5 ml of 2 M HNO₃ and at the end with 20 ml of water. 5 ml of methanol was added on top of the adsorber, and passed through it until the level of methanol reached just the surface of the adsorber. Then water was added to it and connected to the other two columns. A certain volume of water sample was filtered through a 0.45 µm filter and then passed through the three column system, Dowex 1X-8, RP-C18 silica adsorber and Chelex-100, respectively. The columns were then separated. The anion and cation exchanger columns were washed with 10 ml of 2 M HNO₃ and the C₁₈ column with 10 ml of 1 M HCl. The flow rate of eluents was 1 ml min⁻¹. The Ni content of each eluted solution was determined by FAAS.

RESULTS AND DISCUSSION

The treatment of Graphene oxide with covalently linked porphyrin (GO–H₂NP) can lead to the derivatization of both the edge carboxyl and surface hydroxyl functional groups via formation of amides (Karousis et al., 2011) or carbamate esters (Ballester et al., 2000), respectively. The formation of GO–H₂NP was followed by ATR-IR spectroscopy. Initially, in the spectrum of GO, the carbonyl vibration appears at 1716 cm⁻¹, while there are fingerprints at 3616 and 3490 cm⁻¹ due to the presence of hydroxyl species at the basal plane of graphene. The covalent linkage of H₂NP with the acyl chloride activated GO is evident from the presence of a band at 1630 cm⁻¹, which is characteristic for the carbonyl groups of the amide units (Karousis et al., 2011) (Figure 1). The amount of porphyrin attached onto the graphene sheet was evaluated by thermogravimetric analysis. As compared with the TGA results of pure graphite, which is thermally stable up to 900°C under nitrogen, and GO which decomposes above 600°C, after having lost the oxygenated species at 240°C (that is, 14.7% weight loss), the 6% weight loss occurred in the temperature range 250 to 550°C for the GO–H₂NP material, is attributed to

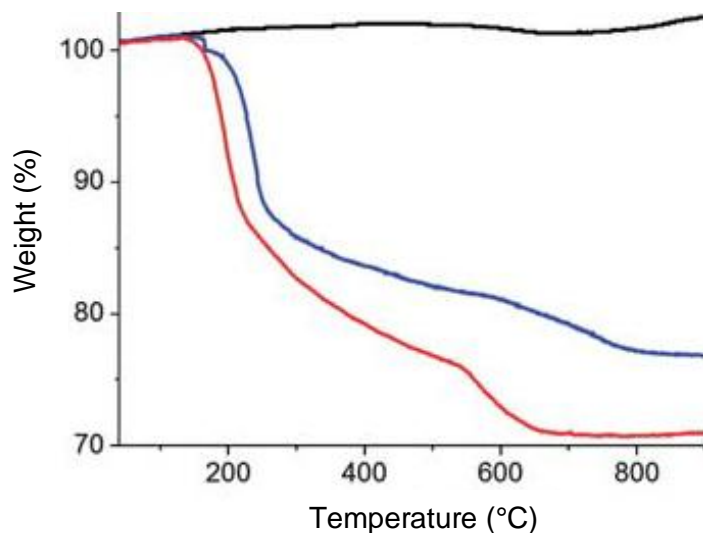


Figure 2. The TGA graphs of graphite (black), GO (blue) and GO-H₂NP (red), obtained under an inert atmosphere.

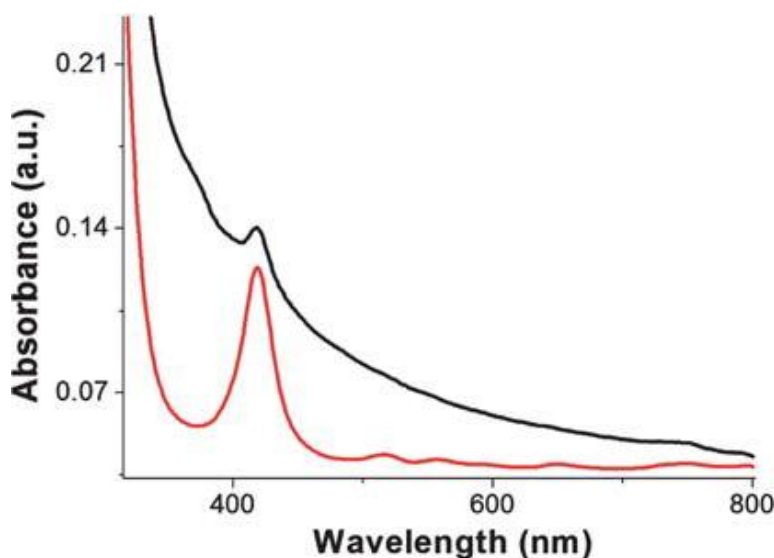


Figure 3. The UV-vis spectra of GO-H₂NP (black) and free H₂NP (red), obtained in DMF.

the decomposition of H₂NP (Figure 2).

The GO-H₂NP material forms a stable dispersion in DMF at a concentration not exceeding 1 mg mL⁻¹. The electronic absorption spectrum of GO-H₂NP in DMF (Figure 3), shows (i) a broad signal monotonically decreasing from the UV to the visible region, which is attributed to GO and (ii) a characteristic band at 420 nm (Soret-band) corresponding to the covalently grafted H₂NP units (the Q-bands at 516, 557, 589 and 648 nm were flattened to the base line in the GO-H₂NP material).

Interestingly, the absorption of porphyrin in the GO-H₂NP material is broadened, shortened and bathochromically shifted (ca. 2 nm) as compared to that of the free H₂NP, a result that corroborates not only the linkage of porphyrin with the GO sheets but also electronic interactions between the two species (that is, GO and H₂NP) in the ground state. These results are in agreement with studies based on other hybrid systems consisting of porphyrins covalently grafted to carbon nanotubes and nanohorns (Karousis et al., 2011).

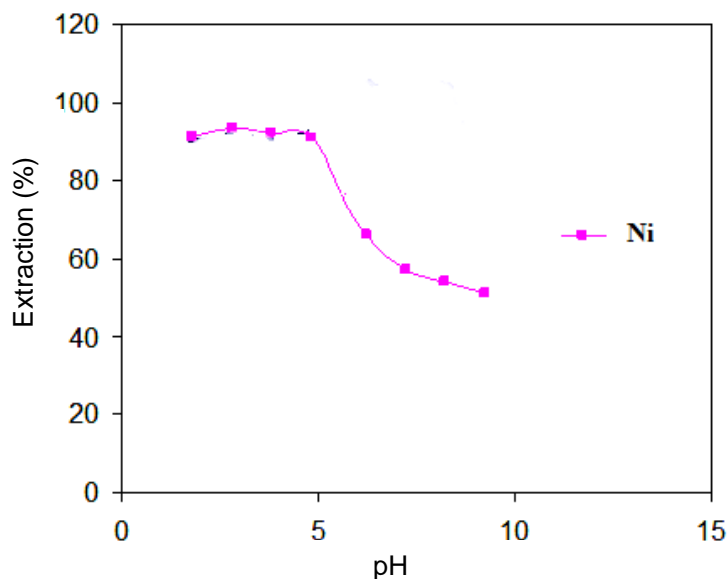


Figure 4. Extraction percentage of Ni (II) against pH.

Stability studies

The stability of the newly synthesized GO-H₂NP phases was performed in different buffer solutions (pH 1, 2, 3, 4, 5, 6 and 0.1M sodium acetate) in order to assess the possible leaching or hydrolysis processes. Because the metal capacity values determined in Section 3.2 revealed that the highest one corresponds to Ni(II), this ion was used to evaluate the stability measurements for the GO-H₂NP phase (Chang et al., 2001). The results of this study proved that the GO-H₂NP is more resistant than the chemically adsorbed analog, especially in 1.0, 5.0 and 10.0 M hydrochloric acid with a hydrolysis percentage of 2.20, 6.10 and 10.50 for each phase, respectively.

Thus, these stability studies indicated the suitability of each phase for application in various acid solutions especially concentrated hydrochloric acid and extension of the experimental range to very strong acidic media which is not suitable for other normal and selective chelating ion exchangers based on a nano polymeric matrix (Manzoori et al., 1998). Finally, the GO-H₂NP phases were also found to be stable over a range of 1 year during the course of this work.

GO-H₂NP is insoluble in water. Primary investigations revealed that surfactant coated C₁₈ could not retain Ni(II) cations, but when modified with the GO-H₂NP, it retains these cations selectively. It was then decided to investigate the capability of the GO-H₂NP as a ligand for simultaneous preconcentration and determination of Ni on admicell. The C₁₈ surface in acidic media (1 < pH < 6) attracts protons and becomes positively charged. The hydrophilic part of SDS (-SO₃⁻), is attached strongly to

these protons. On the other hand, the GO-H₂NP is attached to the hydrophobe part of SDS and retain small quantities of metallic actions (Tao et al., 1998).

Effect of pH

The effect of pH of the aqueous solution on the extraction of 100 ng of each of the cations Ni(II) was studied in the pH range of 1 to 10. The pH of the solution was adjusted by means of either 0.01 mol L⁻¹ HNO₃ or 0.01 mol L⁻¹ NaOH. The results indicate that complete chelation and recovery of Ni (II) occurs in a pH range of 2 to 4 shown in Figure 4. It is probable that at higher pH values, the cations might be hydrolyzed and complete desorption does not occur. Hence, in order to prevent hydrolysis of the cations and also keeping SDS on the C₁₈, pH=2.0 was chosen for further studies.

Effect of flow rates of solutions

The effect of the solution's flow rate of the cations on their chelation on the substrate was also studied. It was indicated that flow rates of 1 to 5 ml min⁻¹ would not affect the retention efficiency of the substrate. Higher flow rates cause incomplete chelation of the cations on the sorbent. The similar range of flow rate for chelation of cations on modified C₁₈ with SDS and a GO-H₂NP has been reported in the literature (Chang et al., 2001). A flow rate of 1 to 2 ml min⁻¹ for desorption of the cations with 4 ml of 4 M HNO₃ has been found suitable. Higher flow rates need a larger volume of acid. Hence, flow rates of 5 ml min⁻¹ and 2 ml min⁻¹ were used for sample solution and

Table 2. Effect of foreign ions on the recovery of 100 ng of Ni.

Diverse ion	Amounts taken (mg) added to 50 ml	% Found	%Recovery of Ni ²⁺ ion
Na ⁺	92.4	1.19(2.6) ^a	98.7(1.9)
K ⁺	92.5	1.38(2.1)	98.9(2.6)
Mg ²⁺	14.5	0.8(1.8)	98.7(1.8)
Ca ²⁺	28.3	1.29(2.0)	99.0(1.8)
Sr ²⁺	3.45	2.85(2.2)	98.2(2.0)
Ba ²⁺	2.66	3.16(2.5)	98.3(2.0)
Mn ²⁺	2.64	1.75(2.3)	98.5(1.8)
Co ²⁺	2.17	1.4(2.6)	98.1(2.2)
Ni ²⁺	2.65	2.0(2.6)	98.4(2.4)
Zn ²⁺	2.74	1.97(2.1)	98.7(2.2)
Cd ²⁺	2.53	1.9(2.0)	98.8(2.8)
Bi ³⁺	2.55	2.7(1.4)	98.7(2.7)
Cu ²⁺	2.46	2.81(2.3)	98.7(2.6)
Fe ³⁺	2.60	3.45(2.4)	96.6(2.8)
Cr ³⁺	1.70	2.92(2.6)	97.3(2.4)
UO ²⁺	2.89	1.3(2.4)	98.3(2.2)
NO ₃ ⁻	5.8	2.3 (2.3)	98.5(2.6)
CH ₃ COO ⁻	5.0	2.2(2.5)	96.6(2.2)
SO ₄ ²⁻	5.0	2.9(3.0)	97.7(2.1)
CO ₃ ²⁻	5.6	1.8(2.5)	99.3(2.5)
PO ₄ ³⁻	2.5	2.1(2.0)	98.9(2.0)

^a:Values in parentheses are CVs based on three individual replicate measurements.

eluting solvents throughout, respectively.

Effect of the GO–H₂NP quantity

To study an optimum quantity of the GO–H₂NP on quantitative extraction of Ni, 50 ml portions of solutions containing 100 ng of each cation were passed through different columns, the sorbent of which were modified with various amounts, between 10 to 50 mg of the GO–H₂NP. The best result was obtained on the sorbent which was modified with 40 mg of the GO–H₂NP.

Figures of merit

The breakthrough volume is of prime importance for solid phase extractions. Hence, the effect of sample volume on the recovery of the cations was studied. 100 ng of each cation was dissolved in 50, 100, 500 and 1000 ml of water. It was indicated that in all the cases, chelation and desorption of the cations were quantitative. It was then concluded that the breakthrough volume could be even more than 1000 ml. Because the sample volume was 1000 ml and the cations were eluted into 10 ml solution, the enrichment factor for both cations is 100, which is

easily achievable. The maximum capacity of 1.5 g of the substrate was determined as follows; 500 ml of a solution containing 50 mg of each cation was passed through the column. The chelated ions were eluted and determined by FAAS.

The maximum capacity of the sorbent for three individual replicates was found to be 20.0 ± 0.9 µg of each cation. The limit of detection (3σ) for the cations (Moghimi and Abdouss 2012) was found to be 5.40 ngL⁻¹ for Ni ions. Reproducibility of the method for extraction and determination of 100 ng of each cation in a 50 ml solution was examined. As the results of seven individual replicate measurements indicated, they were 2.85 and 2.98 % for Ni (II).

Effect of foreign ions

The effect of foreign ions was also investigated on the measurements of Ni. Here a certain amount of foreign ion was added to 50 ml of the sample solution containing 100 ng of each Ni(II) with a pH of 2.5. The amounts of the foreign ions and the percentages of the recovery of Ni are listed in Table 2. As it is observed, it is possible to determine Ni without being affected by the mentioned ions.

Table 3. Recovery of Ni contents of water samples.

Variables	Nickel	Amount added (μg)	Found (μg)	Recovery (%)
Sample distilled water (100 ml)	Ni	-	-	-
		0.050	0.048(3.44) ^a	97
		0.100	0.099(2.90)	98
Tap water (100 ml)	Ni	-	0.018(4.59)	-
		0.050	0.067(3.47)	99
Snow water (50 ml)	Ni	-	0.046(3.60)	-
		0.100	0.150(2.45)	98.0
Rain water (100 ml)	Ni	-	0.038(3.50)	-
		0.100	0.146(2.40)	98
Synthetic sample 1 Na^+ , Ca^{2+} , Fe^{3+} , Co^{2+} , Cr^{3+} , Hg^{2+} , 1 mg L^{-1}	Ni	-	-	-
		0.100	0.102(3.40)	96
Synthetic sample 2 K^+ , Ba^{2+} , Mn^{2+} , Cd^{2+} , Ni^{2+} , Zn^{2+} , 1 mg L^{-1} of each cation	Ni	-	-	-
		0.100	0.101(2.70)	99

^a: Values in parentheses are CVs based on three individual replicate measurements.

Table 4. Results of speciation of Ni in different samples by the three columns system.

Variables	Tap water (1000 ml)	Sample water (1000 ml) ^a	River water (50 ml)
Column	Ni(μg)	Ni(μg)	Ni(μg)
Dowex 1X8	-	-	-
Silica C-18	-	-	-
Chelex-100	0.018(4.5) ^b	0.108(3.1)	0.106(2.6)

^a: This was a solution containing $0.1 \mu\text{g}$ of each cation in 1000 ml of distilled water, ^b: Values in parentheses are CVs based on three replicate analyses. The samples are the same as those mentioned in Table 4.

Analysis of the water samples

The prepared sorbent was used for analysis of real samples. To do this, the amounts of Ni were determined in different water samples namely: distilled water, tap water of Tehran (taken after 10 min operation of the tap), rain water (Tehran, 20 January, 2013), snow water (Tehran, 7 February, 2013), and two synthetic samples containing different cations. The results are tabulated in Table 3. As it is seen, the amounts of Ni added to the water samples are extracted and determined quantitatively which indicates accuracy and precision of the present method.

Separation and speciation of cations were completed by three columns system. It is possible to pre-concentrate and at the same time separate the neutral metal complexes of GO-H₂NP, anionic complexes and free ions from each other by this method (Moghimy and Abdouss 2012). Water samples were passed through the three connected columns: anion exchanger, C₁₈-silica adsorber and chelating cation exchanger. Each species of Ni is retained in one of the columns; anionic complexes

in the first column, neutral complexes of GO-H₂NP in the second, and free ions in the third. The results of passing certain volumes of different water samples through the columns are listed in Table 4.

According to the results, it is indicated that Ni is present only as cations. On the other hand, the t-test comparing the obtained mean values of the present work with those published indicate no significant difference between them. We have proposed a method for the determination and preconcentration of Ni in water samples using surfactant coated C₁₈ impregnated with a GO with covalently linked porphyrin base. The proposed method offers a simple, highly sensitive, accurate and selective method for the determination of trace amounts of Ni(II) in water samples. The method developed was simple, reliable, and precise for determining Ni in water. Also, the proposed method was free of interference compared to conventional procedures to determine Ni (Moghimy and Abdouss 2012; Zargarani et al., 2008; Changlun et al., 2009; Shitong et al., 2009). The method can be successfully applied to the separation and determination of Ni in binary mixtures.

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